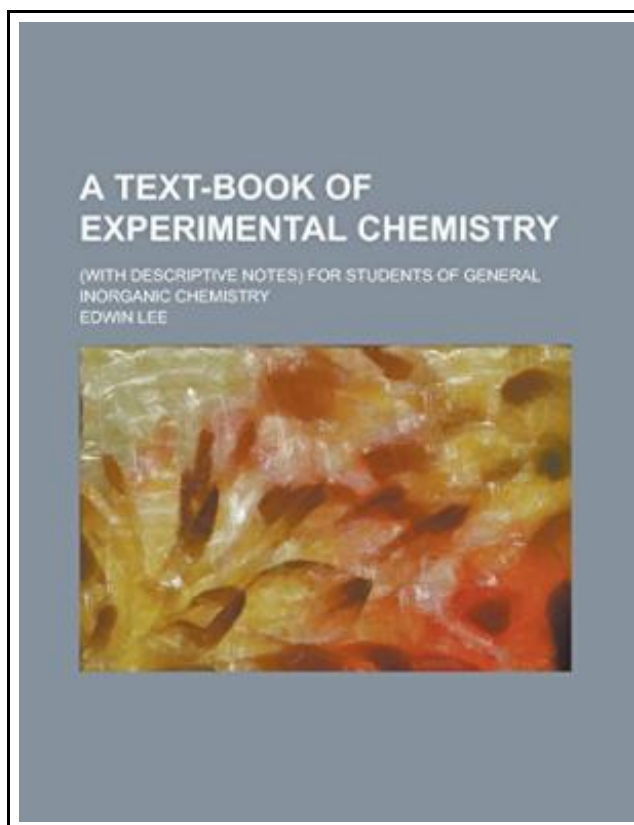


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Reviews




*The publication is great and fantastic. It really is simplistic but surprises within the 50 % from the publication. Your daily life span will be change when you comprehensive reading this article book.
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A TEXT-BOOK OF EXPERIMENTAL CHEMISTRY; (WITH DESCRIPTIVE NOTES) FOR STUDENTS OF GENERAL INORGANIC CHEMISTRY (PAPERBACK)



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Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1908 edition. Excerpt: . (nitrogen peroxide), N_2O , and N_2O_3 . Heat about 8 grams of dry lead nitrate in a test tube. Note the color and odor of the gas evolved. The vapor may be condensed in a U-tube surrounded by a freezing mixture, or a portion of it may be conducted into a concentrated solution of NaOH contained in a test tube. If all of the gas is not absorbed by the solution, test the escaping gas for oxygen. Results? What is the name of the gas? Examine the substance which remains in the ignition tube. What is it? What is the name and formula of the evolved gas? Write equations for all reactions. Keep the NaOH solution for Exp. XVII. Label-it. $(N_2O_2) = -8,125$ cal. (e) Nitrogen pentoxide (nitric anhydride), N_2O_5 , does not exist in the free state. It is usually prepared by distilling a mixture of phosphorus pentoxide, P_2O_5 , and nitric acid. It forms colorless, rhombic prisms which are so unstable that they explode violently when heated quickly. They dissolve readily in water with the disengagement of considerable heat, forming nitric acid, HNO_3 . $2HNO_3 + P_2O_5 \rightarrow N_2O_5 + 2HPO_3$. $(N_2O_3\text{--gas}) = -12,000$ cal. $N_2O_5 + H_2O \rightarrow 2HNO_3$. The series of oxides of nitrogen are an excellent illustration of what law of combination? State the law. Which of the oxides are colored? Colorless? $(H, N_2O_3\text{--liq.}) = 41,500$ cal.; $(H, N_2O_3\text{--Aq.}) = 49,000$ cal. From the foregoing equations it is readily seen that all of the oxides of nitrogen are endothermic...

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